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**Crystallographic data for certain alkaloids. III.** By F. M. LOVELL. *Viriamu Jones Laboratory, University College, Cardiff, Wales.*

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The cell dimensions for these substances (Table 1) were obtained from oscillation and Weissenberg photographs using Cu  $K\alpha$  radiation. Densities were determined by flotation. The maximum error in the numerical data given is of the order of 1%.

**Ephedrine,  $C_{10}H_{15}NO \cdot \frac{1}{2} H_2O$**

Ephedrine is reported to occur in two forms: one is anhydrous and the other contains half a molecule of water of crystallization. Attempts to prepare the anhydrous form were unsuccessful. Oscillation photographs of the hydrated form showed the crystals to be orthorhombic with Laue symmetry  $mmm$ . This, in conjunction with the systematic absences, determines the space group uniquely as  $C222_1$ . The results obtained are compared in Table 1 with those given by Gossner & Neff (1935).

**Hordenine sulphate,  $(C_{10}H_{15}NO)_2 \cdot H_2SO_4 \cdot H_2O$**

Crystals were obtained by recrystallization from methanol. Oscillation photographs showed that the crystals were monoclinic (Laue symmetry  $2/m$ ). Systematic absences indicated an  $a$  glide and a screw axis. The space group is therefore uniquely determined as  $P2_1/a$ . The monoclinic angle was measured from a Weissenberg photograph of the [010] zone of reflexions.

**Homatropine,  $C_{16}H_{21}NO_3$**

Crystals were obtained by recrystallization from ethanol. Oscillation photographs determined the space group uniquely as  $P2_1/c$ . The monoclinic angle was again measured from a Weissenberg photograph.

**Homatropine hydrobromide,  $C_{16}H_{21}NO_3 \cdot HBr$**

Crystals were obtained by recrystallization from water. Oscillation and Weissenberg photographs determined the space group uniquely as  $Pcab$ .

In both homatropine and homatropine hydrobromide the nitrogen atom may be optically active, in which case, since the space groups are centro-symmetric, the unit cells must be assumed to contain equal numbers of L- and D-molecules. The alkaloid homatropine is a synthetic product (Manske & Holmes, 1950) so that a racemate is probable.

**Physostigmine,  $C_{15}H_{21}N_3O_2$**

The crystals obtained by recrystallization from ethanol were in the form of plates. Oscillation photographs showed

the Laue symmetry to be  $mmm$  and the only absences to be screw absences along each axis. The space group is therefore determined uniquely as  $P2_12_12_1$ .

**Narcotine,  $C_{22}H_{27}NO_7$**

Long needle-shaped crystals were obtained by recrystallization of the commercial product from ethanol. As in physostigmine, the Laue symmetry and systematic absences determine the space group uniquely as  $P2_12_12_1$ .

**Quinidine benzenate,  $C_{20}H_{24}NO \cdot \frac{1}{4} C_6H_6$**

Crystals were obtained from a solution of quinidine in benzene. The space group was determined uniquely from oscillation photographs as  $P2_12_12_1$ .

The formula given by Groth (1906-19) is



but the density calculated from this formula and the observed cell edges did not agree with that observed. Also the axial ratios

$$a:b:c = 0.735:1:1.24$$

did not agree with those given by Groth ( $a:b:c = 0.6916:1:1.0054$ ).

When the formula  $C_{20}H_{24}NO \cdot \frac{1}{4} C_6H_6$  was adopted, agreement was found between observed and calculated densities. It is concluded that the material examined differed from that used for the observations quoted by Groth.

No further X-ray work on these substances is contemplated.

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**References**

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Table 1

Compound	$a$ (Å)	$b$ (Å)	$c$ (Å)	$\beta$ (°)	Density (g.cm. <sup>-3</sup> )		$Z$	Space group
					Obs.	Calc.		
Ephedrine	7.41	11.40	24.1	—	1.12	1.13	8	$C222_1$
Ephedrine (Gossner & Neff)	7.41	11.25	24.1	—	1.124	1.15	8	$C222_1$
Hordenine sulphate	24.1	8.29	11.75	103	1.31	1.29	4	$P2_1/a$
Homatropine	14.5	15.1	6.97	100	1.21	1.22	4	$P2_1/c$
Homatropine hydrobromide	10.3	16.4	19.25	—	1.48	1.46	8	$Pcab$
Physostigmine	7.24	14.25	14.50	—	1.20	1.22	4	$P2_12_12_1$
Narcotine	7.90	15.4	32.6	—	1.38	1.38	8	$P2_12_12_1$
Quinidine benzenate	9.42	12.8	15.85	—	1.19	1.19	4	$P2_12_12_1$